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Pentaborate(1-) Salts and a Tetraborate(2-) Salt Derived from C_{2-} or C_{3-} Linked Bis(alkylammonium) Dications: Synthesis, Characterization, and Structural (XRD) Studies

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Abstract: The synthesis of a number of pentaborate(1-) salts from cations arising from N-substituted α,α-, α,β-, and α,γ-diaminoalkanes has been attempted in aqueous solution from B(OH)_3 and the appropriate diammine in a 10:1 ratio. Despite relatively mild work-up conditions the pentaborate(1-) salts prepared were not always as anticipated and the following compounds were isolated in good yield: [Me_2NH(CH_2)_2NHMe_2][B_5O_6(OH)_4]_2 (1), [Et_2NH(CH_2)_2NHEt_2][B_5O_6(OH)_4]_2 (2), [Et_2NH][B_5O_6(OH)_4] (3), [Me_2NH_2][B_5O_6(OH)_4] (4), [Me_2NH(CH_2)_2NHMe_2][B_5O_6(OH)_4]_2 (5), [Et_2NH(CH_2)_2NHEt_2][B_5O_6(OH)_4]_2 (6), [Me_3NCH_2CH=CH_2][B_5O_6(OH)]_4 (7), and [Me_3N(CH_2)_2NMe_3][B_5O_6(OH)_4]_2 0.5H_2O (8). The tetraborate(2-) salt, [Me_3N(CH_2)_2NMe_3][B_4O_5(OH)_4]_2 2B(OH)_3 2H_2O (9) was obtained in moderate yield (41%) from a 3:1 reaction of B(OH)_3 with [Me_3N(CH_2)_2NMe_3](OH)_2. All compounds were characterized by spectroscopy (1H, 11B, 13C NMR and IR) and thermal gravimetric analysis (TGA). BET analysis on materials derived thermally from selected samples (1, 2, 6, 7) all had porosities of < 1 m^2/g, demonstrating that they were non-porous. Single-crystal XRD structures were obtained for 2, 3, 7, 8 and 9 and all contain extensive H-bonded polyborate lattices.

Keywords: BET analysis; borate; diaminoalkane salts; oxidipolyborate; pentaborate(1-); tetraborate(2-); TGA; XRD

1. Introduction

Salts and neutral species containing polyoxidoaborate (polyborate) anions occur as minerals and many synthetic derivatives have now also been prepared [1–5]. These compounds are structurally diverse, with the polyborate anions generally comprised of trigonal planar 3-coordinate and tetrahedral 4-coordinate boron centres, linked via oxygen bridges with as many terminal OH or O^− groups as is necessary to complete the required geometry at boron [6–11]. Some polyborate salts are sold as bulk chemicals for use in the glass/vitreous, fire retardant, and agricultural industries [12–14], whilst other synthetic polyborates have interesting and specialized physical e.g., semiconducting, luminescent, non-linear optical (NLO) properties [5,15]. We are interested in synthetic polyborate chemistry and recently have started to investigate non-metal cation stabilized silicate/borate solutions as potential...
bio-stimulants for agricultural use [12,16,17]. Since choline-like cations and tetraalkylammonium cations are known to structure-direct silicate species towards cubic octasilicate [Si₈O₁₂₊ₙ(OH)₈₋ₙ]ⁿ⁻
anions [18,19] our attention has been directed towards N-substituted aminoalkane and N-substituted α,ω-diaminoalkane cations. This has led us to investigate, as part of a broader parallel study on synthetic polyborate chemistry, the interaction in aqueous solution between B(OH)₃ and cations derived from N-substituted α,ω-diaminoalkanes and the results of this investigation are presented in this manuscript. In particular we report herein the synthesis and characterization of seven new polyborate salts together with single-crystal XRD studies on five of these compounds. Schematic drawings of the pentaborate(1-) and tetraborate(2-) anions present in these salts are given in Figure 1. Aspects of solid-state H-bond interactions, important in stabilizing the solid-state structures, are also discussed.

![Schematic drawings of pentaborate(1-) and tetraborate(2-) anions](image)

**Figure 1.** Schematic drawings of (a) pentaborate(1-) anion observed in 1–8, and (b) tetraborate(2-) anion observed in 9.

### 2. Results and Discussion

#### 2.1. Synthesis

Synthetic strategies to prepare pentaborate(1-) salts partnered with organic cations include the reaction of a free base amine with B(OH)₃ in aqueous solution or use of OH⁻ activated ion-exchange resins with quaternary ammonium halide salts [5]. Schubert and co-workers [19] have successfully applied the former method to α,ω-diaminoalkanes and have isolated a number of pentaborate(1-) salts containing a +2 diammonium cations of composition [H₃N(CH₂)ₓNH₃][B₅O₆(OH)₄]₂ (x = 5, 6, 8-12). In this study, the pentaborate(1-) salts [Me₂NH(CH₂)₂NHMe₂][B₅O₆(OH)₄]₂ (1), [Et₂NH(CH₂)₂NHEt₂][B₅O₆(OH)₄]₂ (2), [Et₂NH][B₅O₆(OH)₄] (3), [Me₂NH][B₅O₆(OH)₄] (4), [Me₂NH(CH₂)₂NHMe₂][B₅O₆(OH)₄]₂ (5), and [Et₂NH(CH₂)₃NHEt₂][B₅O₆(OH)₄]₂ (6) were readily prepared as crystalline solids from aqueous solution in yields of 87–99% by the reaction of 10 equivalents of B(OH)₃ with the diaminoalkane in aqueous solution at room temperature. Compounds 2, 3, 5 and 6 are previously unreported. Compound 1 has recently been prepared serendipitously in low yield (~10%) as a by-product from the reaction of a fluorinated borane CF₃(CF₂)₂CH₂(OLi)CH₂PPh₂BH₃ with [PtCl₂(COD)] in the presence of Me₂N(CH₂)₂NMe₂ [20]. Likewise, 4 has been previously obtained in very low yield (a few crystals) as a by-product of a reaction of B(OH)₃ with B₂(NMe₂)₄ followed by crystallization from water [21]. N,N,N',N'-Tetraalkylaminomethane cations are less stable than the tetrasubstituted diaminoethane or diaminopropane cations in aqueous solution and the formation of the corresponding dialkylammonium pentaborate(1-) salts 3 and 4 in high yields is not unexpected and further demonstrates the ready cleavage of the C-N aminal bonds [22–25]. Further characterization data for 1 and 4 are reported within this manuscript. Scheme 1 contains all the reactions studied in this work.
The pentaborate(1-) salt $\text{[Me}_3\text{NCH}_{2}\text{CH}=\text{CH}_2]\text{[B}_5\text{O}_6\text{(OH)}_4]$ (7) was obtained as the isolated product from crystallization of the crude solid sample obtained from the reaction of $\text{[Me}_3\text{NCH}_{2}\text{CH}_2\text{CH}_2\text{NMe}_3]\text{(OH)}_2$ (from by exchange from $\text{[Me}_3\text{N(CH}_2)_3\text{NMe}_3]\text{I}_2$) with $10\text{B(OH)}_3$ followed by removal of the aqueous solvent on a hot water bath and oven drying (65 °C). This product rather than the expected product, $\text{[Me}_3\text{N(CH}_2)_3\text{NMe}_3]\text{[B}_5\text{O}_6\text{(OH)}_4]^2$ was obtained since the $\text{[Me}_3\text{N(CH}_2)_3\text{NMe}_3]^2+$ cation initially present rapidly undergoes a Hofmann elimination [26,27], with elimination of NMe$_3$, under the workup conditions. Repeating the reaction but maintaining the reaction mixture at room temperature during the workup resulted in the high yield (96%) synthesis of expected product $\text{[Me}_3\text{N(CH}_2)_3\text{NMe}_3]\text{[B}_5\text{O}_6\text{(OH)}_4]^2$ which after recrystallization at room temperature from aqueous solution by slow evaporation gave the compound as its hemihydrate, $\text{[Me}_3\text{N(CH}_2)_3\text{NMe}_3]\text{[B}_5\text{O}_6\text{(OH)}_4]^2\cdot0.5\text{H}_2\text{O}$ (8), as crystals suitable for single-crystal XRD studies.
The pentaborate(1-) salt, \([\text{Me}_3\text{N(CH}_2)_2\text{NMe}_3][\text{B}_5\text{O}_6(\text{OH})_4]_2\), has been previously reported by us from the ion exchange method from \([\text{Me}_3\text{N(CH}_2)_2\text{NMe}_3][\text{B}_2(\text{OH})_2\text{H}_2\text{O}]_2\) and \(\text{B(OH)}_3\) in good (80%) yield [28]. We now report that the related salt, \([\text{Me}_3\text{N(CH}_2)_2\text{NMe}_3][\text{B}_4\text{O}_6(\text{OH})_4][\text{B}_2(\text{OH})_3\text{H}_2\text{O}]_2\) (9), could be obtained in moderate yield (41%) by crystallization of an aqueous solution containing a \([\text{Me}_3\text{N(CH}_2)_2\text{NMe}_3][\text{OH}]_2/\text{B(OH)}_3\) in a ratio of 1:3. Reactions of the other amines or \([\text{Me}_3\text{N(CH}_2)_2\text{NMe}_3][\text{OH}]_2/\text{B(OH)}_3\) with \(\text{B(OH)}_3\) at this 1:3 ratio were not explored within this study. Compound 9 contains a tetraborate(2-) rather than a pentaborate(1-) anion, observed for 1–8, and this anion is drawn schematically in Figure 1(b). Spectroscopic (IR and NMR) and thermal analysis data, including porosity measurements on materials obtained thermally, for these compounds 1–9 are described in Section 2.2 and XRD studies on 2, 3, 7, 8 and 9 are described in Section 2.3.

2.2. Thermal Analysis, Porosity Measurements and Spectroscopic Data

Compounds 1–9 gave satisfactory C, H, N (combustion) elemental analysis data and thermal gravimetric analysis (TGA) data consistent with their formulations. Non-metal cation polyborates have been previously reported to decompose thermally (in air) to afford glassy \(\text{B}_2\text{O}_3\) residues [5,19,29]. This thermal decomposition is usually a two-step processes with the lower temperature step associated with dehydration and cross-linking of hydroxyl groups of the polyborate anions to give condensed polyborate salts. The higher temperature step involves oxidation of the cations to afford the glassy residue. Thus for example, 1, dehydrates to a condensed polyborate, \([\text{Me}_2\text{NH(CH}_2)_2\text{NHMe}_2][\text{B}_{10}\text{O}_{16}]\) in the temperature range 150–300 °C with loss of \(4\text{H}_2\text{O}\) then with oxidation of organic cation to leave residual \(5\text{B}_2\text{O}_3\) at 300–650 °C. Compound 9, which is formulated as a tetraborate(2-) salt with 2 interstitial \(\text{B(OH)}_3\) and 2 interstitial \(\text{H}_2\text{O}\) molecules afforded \(3\text{B}_2\text{O}_3\) by a similar two-step process. Selected samples of 1–6 were either heated to 250 °C for 1h (1, 2, 5, 6) or 600 °C for 1 h (1–3, 6) to generate samples approximating to the ‘condensed’ borate salt or the ‘pyrolysed’ glassy solid that could be used for porosity measurements by use of the Brunauer-Emmett-Teller (BET) method [30]. Measured porosities of the thermally treated samples were in the range 0.06–0.94 m\(^2\)/g indicating that they were all essentially non-porous and in accord with previous work [31,32].

\(^1\text{H}, \, ^{13}\text{C}-\) and \(^{11}\text{B}-\text{NMR}\) spectra were recorded on compounds 1–9 (\(\text{D}_2\text{O}\) solvent) and \(^1\text{H}-\) and \(^{13}\text{C}-\) spectra were in accord with the cations present. Exchangeable NH and B-OH protons were absent as specific signals and it is assumed were observed in the HOD signal present at +4.79 ppm as has been observed previously in related non-metal cation systems [31–33]. The pentaborate(1-) anion, although very stable in the solid-state and is readily templated by many organic cations [34], ‘decomposes’ in aqueous solution as various equilibrium determined borate species are rapidly obtained and influenced by the boron concentration, the temperature and the pH of the aqueous solution [35,36]. However, \(^{11}\text{B}\) spectra of 1–8 were obtained and these samples all showed three peaks centred at 1 (−5) 13 (−35%) and 18 (−60%) ppm in a pattern typical of solutions arising from pentaborate(1-) salts [31–34,37]. The \(^{11}\text{B}\) spectra of 9, originating from a tetraborate(2-) salt, was significantly different with three peaks now at 1.3 (5%), 7.4 (43%) and 11.7 (52%). The upfield shift of the ‘average’ \(^{11}\text{B}\) chemical shift on going from a pentaborate (1-) anion (1–8) to a tetraborate(2-) anion (9) is in accord with the average B/charge ratio changing from 5:1 to 2:1 [34].

IR spectra were obtained on all compounds and strong B-O stretches in the fingerprint region 1450–650 cm\(^{-1}\) are readily observed [38]. A band assigned to an asymmetric B-O stretch associated with the tetrahedral boron centre at ~925 cm\(^{-1}\) and usually diagnostic of a pentaborate(1-) anion [30,39] was clearly evident in samples 1–8. There are many fewer reports of IR spectra of tetraborate(2-) species and some have been tabulated [38] but diagnostic bands have not been reported. It should be noted that the tetraborate(2-) salt 9 also displayed a stretch at 927 cm\(^{-1}\) in the region considered diagnostic for a pentaborate(1-) salt.
2.3. Single-Crystal XRD Studies

Compounds 2, 3, 7, and 8 have been investigated by single-crystal X-ray diffraction studies and pertinent crystallographic data are given in the experimental section, with full details available in the Supplementary Material. These studies confirm their formulation based on spectroscopic, thermal and analytical data described in Section 2.2. All structures were ionic and comprised of the expected organic cations for 2, 3, 8 and 9, or the cation for 7 arising through the Hofmann elimination, with associated polyborate anions.

Compounds 2, 3, 7, and 8 all contain the insular pentaborate(1-) anion, \([\text{B}_5\text{O}_6(\text{OH})_4]^-\). Compound 8 is co-crystallized with 0.5 molecules of \(\text{H}_2\text{O}\). There is some disorder present in the cation and/or co-crystallized molecules all pentaborate(1-) anion structures except 3. Bond lengths and bond angles found for C, N atoms within the organic cations [19,28,37,40,41] are as expected as are the B-O bond lengths, OBO and BOB angles observed for the \([\text{B}_5\text{O}_6(\text{OH})_4]^-\) anion [19,21,28,29,31–35]; for full details see the Supplementary Information. H-bonding interactions within polyborate salts are extremely common and anion-anion interactions between pentaborate(1-) units, forming giant lattices, are well documented and are signaled as a strong driving force [5,29,34] in templating the crystallization of these anions from the Dynamic Combinatorial Library [42–45] arising from \(\text{B(OH)}_3\) under basic conditions [35,36]. The pentaborate (1-) anions in 2, 3, 7, and 8 all form multiple anion-anion interactions and 2 and 3 have additional cation-anion NH…O H-bond interactions. The acceptor sites for the four anion-anion H-bond interactions from each independent pentaborate(1-) anion can be codified \([19,37]\) as \(\alpha,\beta,\alpha,\alpha\) and \(\alpha,\alpha,\alpha,\beta\) (two independent anions) for 2, \(\alpha,\alpha,\alpha,\gamma\) for 3, \(\alpha,\beta,\alpha,\alpha\) for 7, and \(\alpha,\beta,\alpha,\alpha\) and \(\alpha,\alpha,\beta,\alpha\) (two independent anions) for 8. Using Etter nomenclature [46] the H-bond interactions per pentaborate are 3 \(R_2^2(8)\) and 1 \(R_2^2(12)\) for 2, and 3 \(R_2^2(8)\) and 1 \(C(8)\) for 3, 7, and 8. Compound 7 adopts the familiar ‘herringbone’ structure [29,31] whereas 2, 3 and 8 are unique. The structure of 3 (Figure 2) is similar to the ‘stepped brickwall’ structure observed in [4-MepyH, 4-Mepy][\(\text{B}_5\text{O}_6(\text{OH})_4\)] [37]; however the structures are different since 3 has a C(8) interaction involving \(\gamma\) acceptor sites rather than \(\beta\) sites. Views of the anion giant structures in 2 and 8 are given in Figures 3 and 4 respectively.

![Figure 2](image-url)  
**Figure 2.** Giant lattice arising from pentaborate(1-)pentaborate(1-) interactions in 3, viewed along the \(b\) axis. Cations have been omitted for clarity, but are situated within ‘cavities’. Within this ‘plane’ each pentaborate(1-) in shown H-bonding to 3 other pentaborate(1-) anions by \(R_2^2(8)\) interactions. A fourth C(8) interaction (not shown) links these planes.
positions with site occupancies of 47.4%, 28.7% and 23.9% and one of the two co-crystallized B(OH)3 units occupying 2 positions with site occupancies of 76% and 24%. A drawing of the atomic labelling found with non-metal cation [40,49–52] or transition metal cation salts [53–56].

The structural characterization of 9 is considered in more detail in this manuscript since it contains a tetraborate(2-) anion and there are far fewer crystallographic reports in the literature featuring this anion; the anion is found partnered with 2Na+ in the well-known borax [47,48] but is less commonly found with non-metal cation [40,49–52] or transition metal cation salts [53–56].

The tetraborate(2-) anion in 9 is found co-crystallized with 2B(OH)3 and 2H2O and the dication [Me3N(CH2)2NMe3]2+. The structure is disordered with the cation occupying three closely related positions with site occupancies of 47.4%, 28.7% and 23.9% and one of the two co-crystallized B(OH)3 units occupying 2 positions with site occupancies of 76% and 24%. A drawing of the atomic labelling
for the components of highest site occupancy is given in Figure 5. Bond lengths and bond angles found for C, N atoms within the [Me₃N(CH₂)₂NMe₃]²⁺ cations are as expected and similar to those reported for the related pentaborate(1-) salt, [Me₃N(CH₂)₂NMe₃][B₅O₆(OH)₁₂] [28]. The B-O bond lengths, OBO and BOB angles observed for the [B₄O₅(OH)₄]²⁻ and B(OH)₃ motifs are in accord with previous tetraborate(2-) [47–56] and polyborates with co-crystallized B(OH)₃ [32,33,39,55,57,58]; for full details see the Supplementary Information.

The tetraborate(2-) anion has four H-bond donor sites and nine potential H bond acceptor sites, and the B(OH)₃ molecules each have three donor and three potential acceptor sites. Again, there are numerous H-bond interactions within the structure and a giant supramolecular anionic H-bonded lattice is formed, with cations situated within the cavities. However, there are no direct anion-anion interactions within the lattice and the B(OH)₃ units serve to bridge tetraborate(2-) anions by acting as ‘spacers’ to expand the lattice so that it can accommodate the relatively large cation, [Me₃N(CH₂)₂NMe₃]. This spacer role for B(OH)₃ has been observed before [32,33,39,55–58]. With reference to Figure 6 it can be seen that the H-bond structure can be envisaged as ‘horizontal’ chains of alternating [B₄O₅(OH)₄]²⁻/B(OH)₃ units held together by R₂²(8) interaction, crosslinked by ‘vertical’ C₂²(8) chains (involving O8,H8…O12,H12…O5,B1,O2,B3) in a regular 2D arrangement. These planes are further linked by additional R₂²(8) interactions into a 3D lattice. Further details of the interactions around the tetraborate(2-) anion is shown in Figure 7 and full details are available in the Supplementary Information. The O1 site is a double H-bond acceptor in two R₂²(8) interactions: O21H21…O1 and O13H13…O1. The H₂O of crystallization further help to H-bond the structure together. The cation is unable to get involved with H-bonding.
Figure 6. A 2D plane in 9 comprised of alternating B(OH)$_3$/[B$_4$O$_5$(OH)$_4$]$^{2-}$ ‘horizontal’ chains and ‘vertical’ C$_{22}^2$(8) crosslinking.

Figure 7. Drawing of the H-bond interactions around the tetraborate(2-) anion in 9, showing three R$_2^2$(8) interactions and parts of the C$_{22}^2$(8) chain.

3. Materials and Methods

3.1. General

All chemicals were obtained commercially. TGA/DSC analysis (in air) were undertaken on an SDT Q600 V4.1 Build 59 instrument (TA Instruments, New Castle, DE, USA) using Al$_2$O$_3$ crucibles, between 10–800 °C (ramp temperature rate of 10 °C min$^{-1}$). FTIR spectra were obtained (KBr pellets) on a Perkin-Elmer 100 FTIR spectrometer (Perkin-Elmer, Seer Green, UK). NMR spectra were obtained on a Bruker Avance 400 spectrometer (Bruker, Coventry, UK) and reported in ppm with positive chemical shifts (δ) to high frequency (downfield) of TMS (1H, 13C) and BF$_3$ OEt$_2$ (11B). BET analysis were performed on a Gemini 2375 analyser (Micromeritics Instrument Corporation, Norcross, GA, USA) with N$_2$ gas as the adsorbent. CHN analysis were obtained from OEA Laboratories Ltd. (Callington, UK). See Section 3.11 for single-crystal XRD methods.
3.2. Synthesis, Spectroscopic, and Analytical Data for 1

B(OH)₃ (6.2 g; 100.2 mmol) was dissolved in H₂O (100 mL). To this solution was added Me₂N(CH₂)₂NMₑ₂ (1.18 g; 10.0 mmol) and the reaction mixture was stirred for 1 h. The solvent was partially removed by rotary evaporation to give white crystals. The crude product of 1 (5.5 g; 99%) was obtained after oven drying at 60 °C for 24 h. C₆H₂B₁₀N₂O₂₀. Anal. Calc.: C = 13.0%, H = 4.7%, N = 5.1%. Found: C = 13.1%, H = 4.7%, N = 5.0%. TGA: 150–300 °C, condensation of pentaborate units with loss of 4H₂O · 4.0% (13.0% calc.); 300–650 °C, oxidation of organic cation to leave residual 5B₂O₃ 62.5% (62.8% calc.). BET /m²·g⁻¹: condensed 0.94, pyrolysed 0.61. NMR. ¹H/ppm: 2.47 (12H, s), 2.90 (4H, s). ¹¹B: 1.1 (14%), 12.9 (29%), 19.2 (57%). ¹³C: 43.4 (CH₃) 53.15 (CH₂). IR (KBr/cm⁻¹): 3412 s, 1424 s, 1300 s, 968 m, 928 m, 792 s, 706 m.

3.3. Synthesis, Spectroscopic, Analytical, and Crystallographic Data for 2

Following the method as described for 1, B(OH)₃ (6.2 g; 100.2 mmol) and Et₂N(CH₂)₂NMe₂ (1.74 g; 10.1 mmol) gave crude product of 2 (5.6 g; 91%). C₁₀H₃₄B₁₀N₂O₂₀. Anal. Calc.: C = 19.6%, H = 5.6%, N = 4.6%. Found: C = 19.7%, H = 5.6%, N = 4.3%. Crystals suitable for single-crystal XRD studies were obtained by recrystallization of the crude product from H₂O. TGA: 100–240 °C, condensation of pentaborate units with loss of 4H₂O 12.0% (11.8% calc.); 240–650 °C, oxidation of organic cation to leave residual 5B₂O₃ 54.6% (53.3% calc.). BET/m²·g⁻¹: condensed 0.43, pyrolysed 0.36. NMR. ¹H/ppm: 1.06 (12H, t), 2.90 (4H, q), 3.08 (4H, s). ¹¹B: 1.1 (20%), 13.4 (30%), 18.6 (50%). ¹³C: 8.76 (CH₃), 46.50 (4CH₂), 47.28 (2CH₂). IR (KBr/cm⁻¹): 3445 s, 3143 s, 1316 m, 1230 s, 1101 s, 926 s, 783 m, 779 m. XRD crystallographic data: C₁₀H₃₄B₁₀N₂O₂₀, Mᵣ = 610.49, Triclinic, P–1, a = 8.3998(5) Å, b = 9.1406(7) Å, c = 18.2066(13) Å, α = 78.439(6)°, β = 86.810(5)°, γ = 88.118(6)°, V = 1367.07(17) Å³, T = 100(2) K, Z = 2, μ(MoKα) = 0.130 mm⁻¹, 11455 reflections measured, 6146 unique (Rᵣₑₐₓ = 0.0461) which were used in all calculations. The final wR₂ was 0.1850 (all data) and R₁ was 0.0748 (I > 2σ(I)).

3.4. Synthesis, Spectroscopic, Analytical, and Crystallographic Data for 3

Following the method as described for 1, B(OH)₃ (1.0 g; 16.2 mmol) and Et₂NCH₂NEt₂ (0.26 g; 1.62 mmol) gave crude product of 3 (0.9 g; 99%). C₄H₁₆B₅N₂O₁₀. Anal. Calc.: C = 16.4%, H = 5.5%, N = 4.8%. Found: C = 16.7%, H = 5.6%, N = 4.7%. Crystals suitable for single-crystal XRD studies were obtained by recrystallization of the crude product from H₂O. TGA: 100–240 °C, condensation of pentaborate units with loss of 2H₂O 12.3% (14.3% calc.); 240–600 °C, oxidation of organic cation to leave residual 2.5B₂O₃ 58.3% (59.6% calc.). BET/m²·g⁻¹: pyrolysed 0.06. NMR. ¹H/ppm: 1.16 (6H, t), 2.95 (4H, q). ¹¹B: 1.2 (20%), 12.9 (25%), 18.3 (55%). ¹³C: 10.45 (CH₃) 42.18 (CH₂). IR (KBr/cm⁻¹): 3436 s, 3377 s, 3228 m, 2977 m, 1435 s, 1361 s, 1251 s, 1103 s, 1026 s, 925 s, 882 m, 697 m. XRD crystallographic data: C₄H₁₆B₅N₂O₁₀, Mᵣ = 292.23, Triclinic, P–1, a = 8.3793(3) Å, b = 8.8187(4) Å, c = 10.0877(4) Å, α = 80.961(3)°, β = 76.057(3)°, γ = 68.269(4)°, V = 670.14(5) Å³, T = 100(2) K, Z = 2, μ(MoKα) = 0.129 mm⁻¹, 5980 reflections measured, 3032 unique (Rᵣₑₐₓ = 0.0144) which were used in all calculations. The final wR₂ was 0.1024 (all data) and R₁ was 0.0363 (I > 2σ(I)).

3.5. Synthesis, Spectroscopic and Analytical Data for 4

Following the method as described for 1, B(OH)₃ (1.0 g; 16.2 mmol) and Me₂NCH₂NMₑ₂ (0.17 g; 1.62 mmol) gave crude product of 4 (0.78 g; 92%). C₃H₁₂B₅N₂O₁₀. Anal. Calc.: C = 9.1%, H = 4.6%, N = 5.3%. Found: C = 9.4%, H = 4.6%, N = 5.3%. TGA: 100–240 °C, condensation of pentaborate units with loss of 2H₂O 14.0% (13.6% calc.); 240–600 °C, oxidation of organic cation to leave residual 2.5B₂O₃ 65.6% (65.8% calc.). NMR. ¹H/ppm: 2.59 (6H). ¹¹B: 1.1 (15%), 13.2 (30%), 18.8 (55%). ¹³C: 34.42 (CH₃). IR (KBr/cm⁻¹): 3437 s, 3380 s, 3053 m, 1441 s, 1460 s, 1252 m, 1103 s, 1024 s, 925 s, 783 m, 765 m.
3.6. Synthesis, Spectroscopic and Analytical Data for 5

Following the method as described for 1, B(OH)₃ (6.2 g; 100.2 mmol) and Me₂N(CH₂)₃NMe₂ (1.30 g; 10.0 mmol) gave crude product of 5 (5.0 g; 88%). C₇H₃₈B₁₀N₂O₂₀. Anal. Calc.: C = 14.8%, H = 5.0%, N = 4.9%. Found: C = 14.7%, H = 5.1%, N = 4.8%. TGA: 100–230 °C, condensation of pentaborate units with loss of 4H₂O 13.2% (12.6% calc.); 230–630 °C, oxidation of organic cation to leave residual 5B₂O₃ 60.2% (61.2% calc.). BET/m² g⁻¹: condensed 0.62. NMR. ¹H/ppm: 3.03 (4H, t), 2.70 (12H, s), 2.05 (2H, m). ¹¹B: 1.1 (20%), 13.2 (25%), 18.8 (55%). ¹³C: 42.8 (4CH₃) 54.14 (2CH₂), 19.9 (CH₂). IR (KBr/cm⁻¹): 3307 s, 2859 s, 1440 s, 1399 s, 1253 m, 1180 m, 1025 s, 924 s, 779 m.

3.7. Synthesis, Spectroscopic and Analytical Data for 6

Following the method as described for 1, B(OH)₃ (6.2 g; 100.2 mmol) and Et₂N(CH₂)₂NEt₂ (1.86 g; 10.0 mmol) gave crude product of 6 (5.4 g; 87%). C₁₁H₃₆B₁₀N₂O₂₀. Anal. Calc.: C = 21.2%, H = 5.8%, N = 4.5%. Found: C = 21.4%, H = 5.9%, N = 4.4%. TGA: 100–230 °C, condensation of pentaborate units with loss of 4H₂O 12.1% (11.5% calc.); 230–630 °C, oxidation of organic cation to leave residual 5B₂O₃ 55.3% (55.7% calc.). BET/m² g⁻¹: condensed 0.43, pyrolysed 0.71. NMR. ¹H/ppm: 1.13 (12H, t), 1.97 (2H, m), 3.05 (12H, m). ¹¹B: 1.1 (25%), 12.7 (30%), 18.2 (45%). ¹³C: 43.4 (CH₃) 53.15 (CH₂). IR (KBr/cm⁻¹): 3412 s, 1424 s, 1300 s, 968 m, 928 m.

3.8. Synthesis, Spectroscopic, Analytical, and Crystallographic Data for 7

Me₂N(CH₂)₃NMe₂ (2.6 g, 20 mmol) dissolved in CH₂CN (50 mL) was refluxed with MeI (11.4 g, 80 mmol) for 4 h. The white solid formed, [Me₂N(CH₂)₃NMe₃]₂ (8.28 g, 100%) was isolated by filtration, washed with Et₂O and used without further purification. [NMR. ¹H/ppm: 2.31 (2H, quin.), 3.13 (18H, s), 3.37 (4H, t). ¹³C: 17.41 (CH₂), 53.25 (CH₃), 62.39 (CH₂). IR (KBr/cm⁻¹): 3447 s, 3011 s, 2957 m, 1624 m, 1486 s, 1475 s, 1408 m, 1244 m, 1054 m, 973 s, 944 s, 901 s, 762 m, 545 m.] [Me₂N(CH₂)₃NMe₂]₂ (0.67 g, 1.62 mmol) was dissolved in H₂O (20 mL) and stirred with excess (13.5 g) of DOWEX 550A ion-exchange resin (OH⁻ form) for 24 h. The slurry was filtered and B(OH)₃ (1.0 g, 16.2 mmol) was added to the filtrate. The solution was left for 4 h before removal of solvent under reduced pressure at 85 °C to yield a ‘damp’ solid which was dried in an oven at 75 °C for 3 h to yield a white solid (0.93 g). ¹¹NMR analysis showed this to be a ~ 2:1 mixture of 7 and 8. Recrystallization of 0.2 g of this solid resulted in 0.1 g of 7 with crystals suitable for sc-XRD analysis. C₆H₁₈B₅NO₁₀. Anal. Calc.: C = 22.6%, H = 5.7%, N = 4.4%. Found: C = 23.0%, H = 5.7%, N = 4.4%. TGA: 100–285 °C, condensation of pentaborate units with loss of 2H₂O 11.7% (11.3% calc.); 285–700 °C, oxidation of organic cation to leave residual 2.5B₂O₅ 55.2% (54.7% calc.). NMR. ¹H/ppm: 3.00 (9H, s), 3.85 (2H, d), 5.63 (2H, dd), 5.96 (1H, m). ¹¹B: 1.1 (1%), 13.4 (17%), 17.1 (82%). ¹³C: 52.30 (CH₃), 68.39 (CH₂), 124.43 (CH), 129.14 (CH₂). IR (KBr/cm⁻¹): 3437 s, 3262 m, 1471 m, 1427 s, 1414 s, 1389 s, 1311 s, 1168 m, 1105 s, 1012 s, 921 s, 777 s, 724 m, 708 s. XRD crystallographic data: C₆H₁₈B₅NO₁₀, Mr = 318.26, monoclinic, P₂₁/c, a = 9.54050(10) Å, b = 16.1031(2) Å, c = 9.43280(10) Å, α = γ = 90°, β = 90.1710(10)°, V = 1449.17(3) Å³, T = 100(2) K, Z = 4, Z' = 1, μ(CuKα) = 1.096 mm⁻¹, 13531 reflections measured, 2661 unique (Rint = 0.0236) which were used in calculations. The final wR₂ was 0.0845 (all data) and R₁ was 0.0314 (I > 2σ(I)).

3.9. Synthesis, Spectroscopic, Analytical, and Crystallographic Data for 8

[Me₂N(CH₂)₃NMe₃]₂ (0.40 g, 0.97 mmol), prepared as described for 7, was dissolved in H₂O (20 mL) and stirred with excess (8.0 g) of DOWEX 550A ion-exchange resin (OH⁻ form) for 24 h. The slurry was filtered and B(OH)₃ (0.6 g, 9.7 mmol) was added to the filtrate. Partial evaporation of the solution afforded a white crystalline solid (0.57 g, 96%), separated by filtration. These crystals were suitable for sc-XRD studies. C₆H₁₃B₁₀N₂O₂₁. Anal. Calc.: C = 17.9%, H = 5.5%, N = 4.6%. Found: C = 18.2%, H = 5.4%, N = 4.6%. TGA: 250–300 °C, condensation of pentaborate units with loss of 4.5 H₂O 13.0% (13.4% calc.); 300–700 °C, oxidation of organic cation to leave residual 5B₂O₃ 57.3% (57.5% calc.). NMR. ¹H/ppm: 2.29 (2H, quint.), 3.11 (18H, s), 3.33 (4H, t). ¹¹B: 1.1 (3%), 13.1 (35%), 18.0 (62%).
\[ V = \gamma \mu \theta \]
Supplementary Materials: The following are available online at http://www.mdpi.com/1420-3049/25/1/53/s1, Single crystal XRD data, including Figures of the structures with numerous H-bond interactions.

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Sample Availability: Samples of the compounds are not available from the authors.